

Modified Enlarged 24pt
OXFORD CAMBRIDGE AND RSA EXAMINATIONS

Tuesday 06 October 2020 – Afternoon

AS Level Chemistry A

H032/01 Breadth in chemistry

Time allowed: 1 hour 30 minutes
plus your additional time allowance

YOU MUST HAVE:

the Data Sheet for Chemistry A

YOU CAN USE:

a scientific or graphical calculator
an HB pencil

Please write clearly in black ink.

Centre number

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Candidate number

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First name(s) _____

Last name _____

READ INSTRUCTIONS OVERLEAF



INSTRUCTIONS

Use black ink. You can use an HB pencil, but only for graphs and diagrams.

Write your answer to each question in the space provided. If you need extra space use the lined pages at the end of this booklet. The question numbers must be clearly shown.

Answer ALL the questions.

Where appropriate, your answer should be supported with working. Marks might be given for using a correct method, even if your answer is wrong.

INFORMATION

The total mark for this paper is 70.

The marks for each question are shown in brackets [].

ADVICE

Read each question carefully before you start your answer.

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SECTION A

You should spend a maximum of 25 minutes plus your additional time allowance on this section.

Answer ALL the questions.

Write your answer to each question in the box provided.

1 Which substance contains polar molecules? [1]



Your answer

2 What is the formula of silver carbonate?
[1]



Your answer

3 Which statement explains why ice is less dense than water? [1]

- A Hydrogen bonds are stronger in ice than in water.**
- B Hydrogen bonds hold H₂O molecules apart in ice.**
- C Ice is a solid but water is a liquid.**
- D Ice contains hydrogen bonds, but water does not contain hydrogen bonds.**

Your answer

- 4 Some Group 2 compounds can be used to neutralise acid soils and to treat acid indigestion.

Which Group 2 compound would NOT be suitable for either use? [1]



Your answer

5 Which p-block element contains atoms with one unpaired electron? [1]

A Al

B Si

C P

D S

Your answer

- 6 The equation for a redox reaction is shown below.



Which statement is correct? [1]

- A Cl is both oxidised and reduced.
- B Cl is oxidised and O is reduced.
- C O is both oxidised and reduced.
- D O is oxidised and Cl is reduced.

Your answer

☐

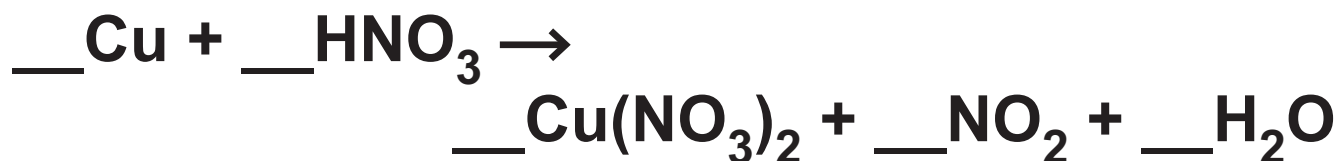
- 7 Potassium ferrate(VI) contains two potassium ions for every ferrate(VI) ion.

What is the formula of the ferrate(VI) ion? [1]



Your answer

8 The unbalanced equation for the reaction of copper with concentrated nitric acid is shown below.



What is the number of moles of HNO_3 that react with 1 mole of Cu? [1]

A 2

B 3

C 4

D 6

Your answer

- 9 2.0 dm^3 of Cl_2 gas reacts with 2.0 dm^3 of ClF_3 gas to form 6.0 dm^3 of a gaseous compound.

The reaction has 100% atom economy and all volumes are measured at the same temperature and pressure.

What is the molecular formula of the compound formed? [1]

A ClF

B Cl_2F_3

C Cl_3F_2

D Cl_3F_3

Your answer

10 Which sample contains the greatest number of molecules? [1]

A 1 g of methanol, CH_3OH

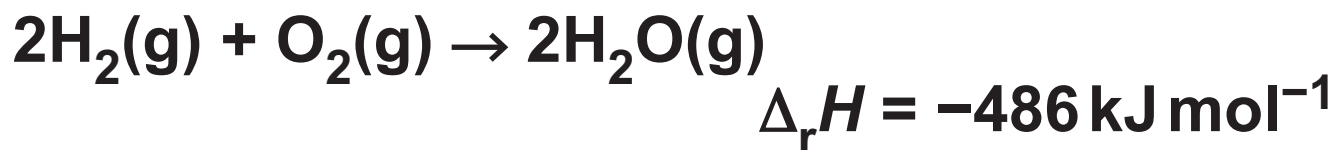
B 2 g of nitrogen dioxide, NO_2

C 3 g of phosphorus, P_4

D 4 g of iodine, I_2

Your answer

11 Hydrogen and oxygen react as shown below.



Bond enthalpies are shown in the table.

Bond	H–H	O=O
Bond enthalpy / kJ mol ⁻¹	+436	+498

What is the bond enthalpy, in kJ mol⁻¹, for the O–H bond? [1]

A +221

B +355

C +464

D +928

Your answer

12 Hydrogen gas can be produced as shown below.

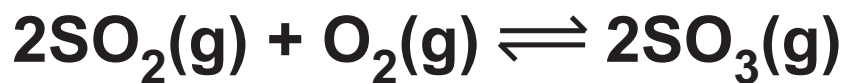


Which conditions produce the greatest equilibrium yield of hydrogen? [1]

- A Low temperature and high pressure**
- B Low temperature and low pressure**
- C High temperature and high pressure**
- D High temperature and low pressure**

Your answer

13 The reversible reaction below is in equilibrium.



The equilibrium concentrations are shown in the table.

Substance	SO₂(g)	O₂(g)	SO₃(g)
Equilibrium concentration / mol dm⁻³	4.00	2.40	1.44

What is the numerical value of K_c ? [1]

A 0.0375

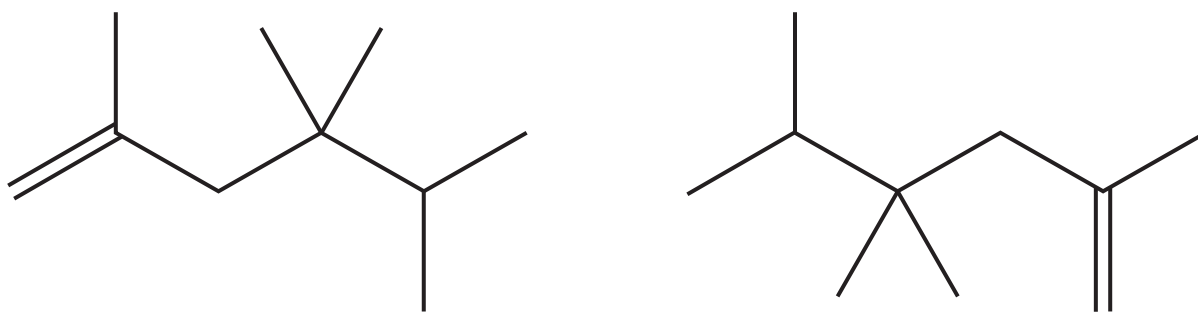
B 0.0540

C 0.150

D 18.5

Your answer

14 Which statement is correct for the two structures below? [1]



- A They have the same empirical formula.**
- B They have different relative molecular masses.**
- C They are structural isomers.**
- D They have different functional groups.**

Your answer

☐

15 Which property explains the low reactivity of alkanes? [1]

- A Electron pair repulsion between σ -bonds**
- B Free rotation about σ -bonds**
- C High C–C bond enthalpy**
- D High polarity of the C–H bonds**

Your answer

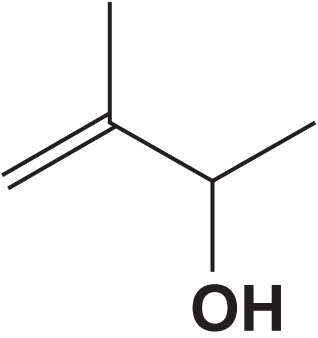
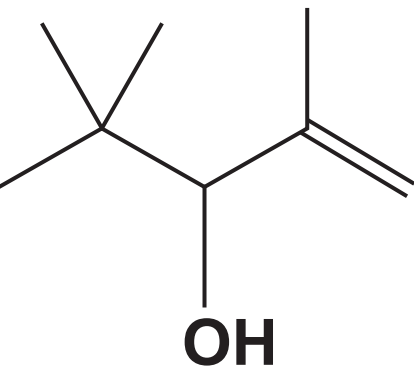
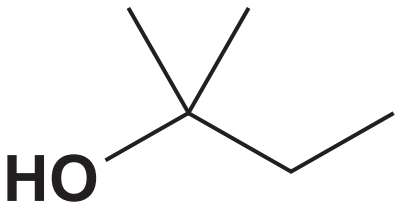
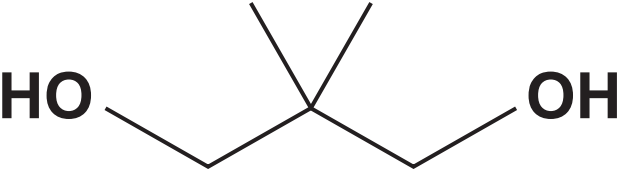
16 What are the correct reagents for the conversion below? [1]



- A Br₂ and H₂SO₄**
- B Br₂ and NaOH**
- C NaBr and H₂SO₄**
- D NaBr and NaOH**

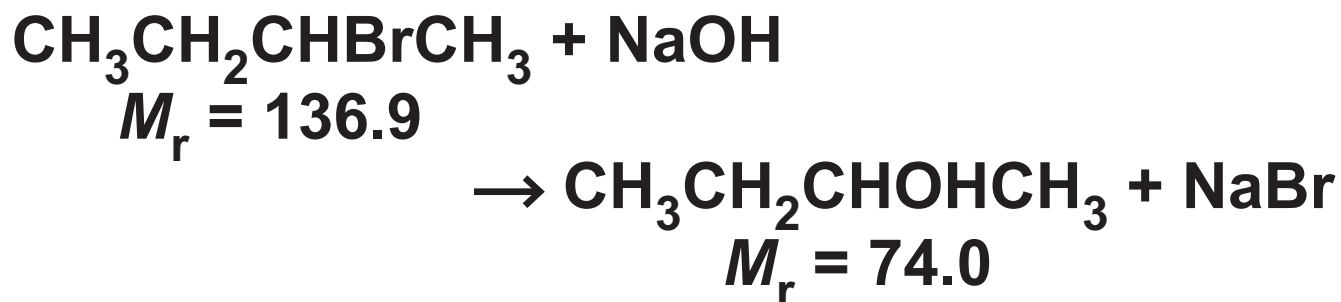
Your answer

17 Which compound could react with BOTH $K_2Cr_2O_7/H_2SO_4$ in an oxidation reaction AND an acid catalyst (e.g. H_2SO_4) in an elimination reaction? [1]

A	
B	
C	
D	

Your answer

18 A student reacts 24.24 g of 2-bromobutane in the reaction below.



The reaction produces 4.81 g of $\text{CH}_3\text{CH}_2\text{CHOHCH}_3$.

What is the percentage yield of $\text{CH}_3\text{CH}_2\text{CHOHCH}_3$? [1]

A 10.7%

B 19.8%

C 36.7%

D 54.1%

Your answer

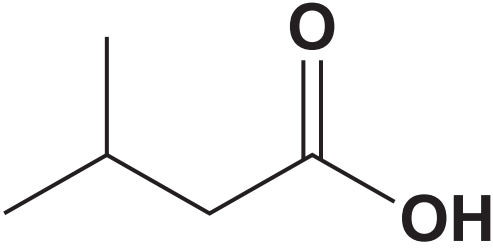
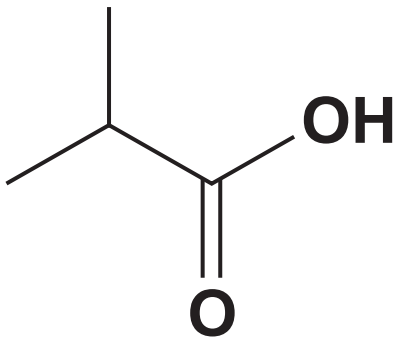
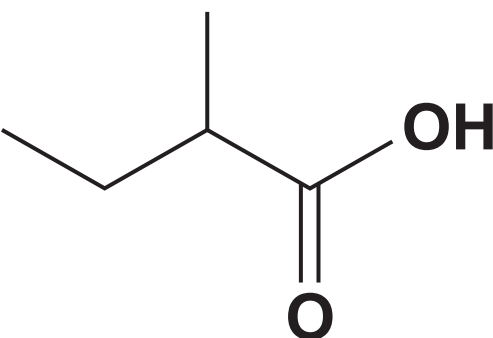
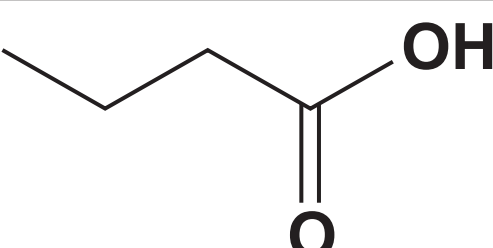
19 Which row describes a nucleophile? [1]

A	electron pair donor	attracted to high electron density
B	electron pair donor	attracted to low electron density
C	electron pair acceptor	attracted to high electron density
D	electron pair acceptor	attracted to low electron density

Your answer

20 The mass spectrum of a carboxylic acid contains peaks at $m/z = 29$ and $m/z = 102$.

Which compound could have produced the spectrum? [1]

A	
B	
C	
D	

Your answer

SECTION B

Answer ALL the questions.

21 This question is about atoms, isotopes and mass spectrometry.

(a) Complete the table to show the number of electrons that can fill the first four shells. [1]

Shell	1st shell	2nd shell	3rd shell	4th shell
Number of electrons				

(b) Most elements contain atoms of different isotopes.

State any differences and similarities between the atomic structures of isotopes of the same element.

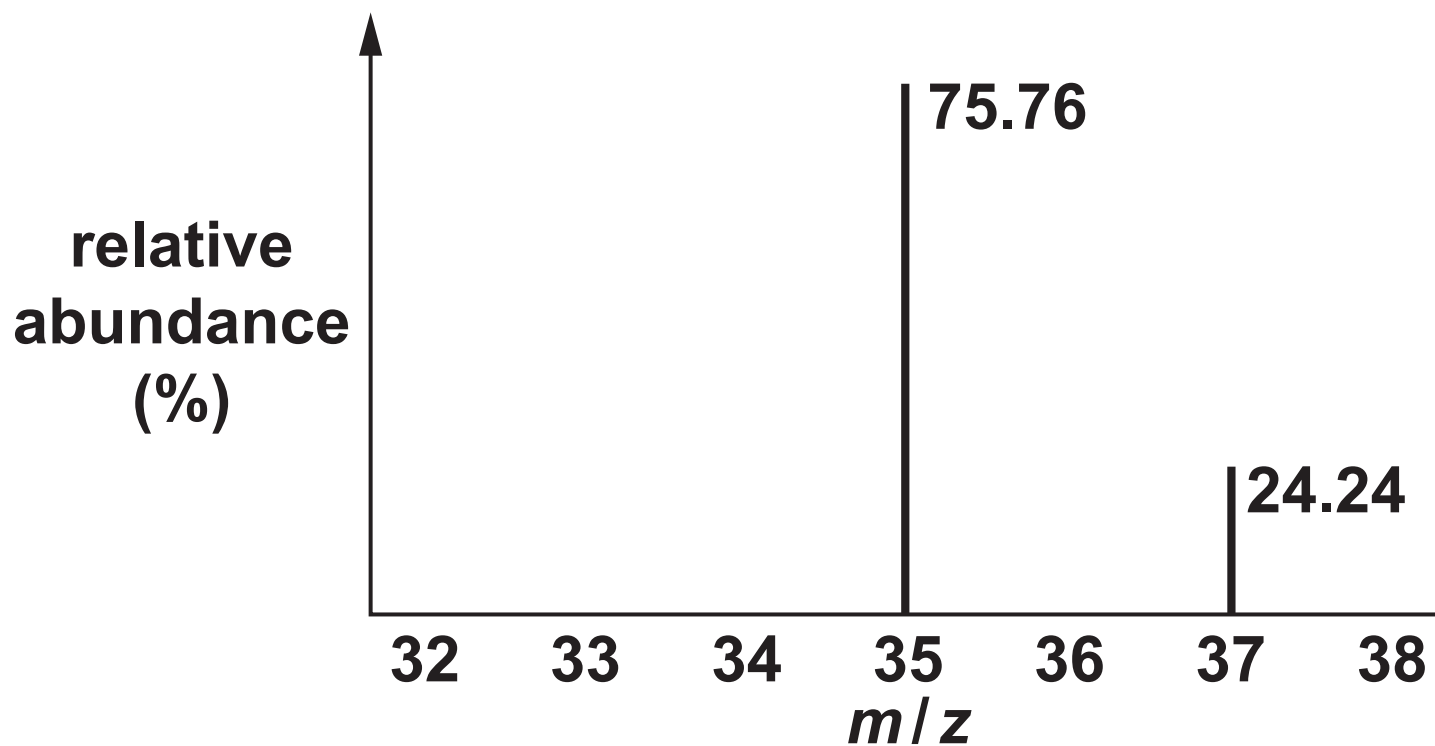
Differences _____

Similarities _____

[2]

(c) Mass spectrometry can be used to identify the isotopes of chlorine.

Part of the mass spectrum of chlorine is shown below.



(i) Calculate the relative atomic mass of chlorine.

Give your answer to 2 decimal places.

relative atomic mass = _____ [2]

- (ii) The mass spectrum of chlorine, Cl_2 , also contains three molecular ion, M^+ , peaks. One of the M^+ peaks has an m/z value of 72.

Suggest why an M^+ peak at $m/z = 72$ is observed and predict the m/z values of the other two M^+ peaks.

Peak at $m/z = 72$ _____

m/z values of the other two

M^+ peaks _____

[2]

22 This question is about compounds of bromine.

(a) Bromine reacts with phosphorus, P_4 , to form phosphorus tribromide, PBr_3 .

(i) Complete the electron configuration of a bromine atom.

$1s^2$ _____ [1]

(ii) Write the equation for the reaction of phosphorus with bromine.

_____ [1]

- (b) A compound of bromine is a solid at room temperature. The electrical conductivity of the compound at different physical states is shown in the table.

Physical state	Electrical conductivity
solid	poor
liquid	good

Name the type of lattice in the compound at room temperature and explain the different electrical conductivities.

Name of lattice _____

Explanation for different conductivities _____

[2]

(c) Bromine reacts with fluorine to form compound A.

Compound A is a liquid at room temperature and pressure but can easily be vaporised.

When vaporised, 0.428 g of A produces 76.0 cm³ of gas at 1.00×10^5 Pa and 100 °C.

Determine the molar mass and molecular formula of compound A.

molar mass of A = _____ g mol⁻¹

molecular formula of A = _____

[5]

23 This question is about barium hydroxide.

(a) Barium hydroxide is an alkali which releases hydroxide ions, OH^- , in aqueous solution.

A barium hydroxide solution contains 3.89 g of $\text{Ba}(\text{OH})_2$ in 100 cm^3 at 20°C .

Calculate the concentration of hydroxide ions, OH^- , in mol dm^{-3} , of this solution at 20°C .

Give your answer to 3 significant figures.

**concentration
of OH^- ions = _____ mol dm^{-3} [3]**

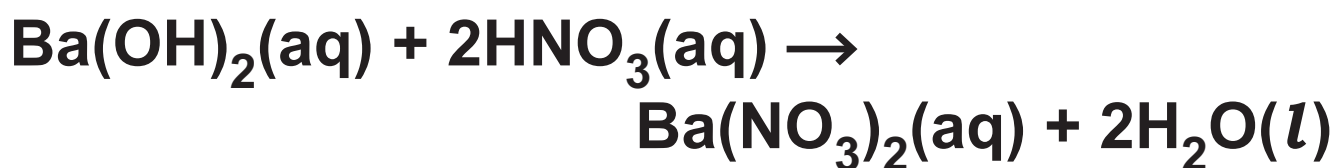
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- (b) A student carries out a titration to determine the concentration of an aqueous solution of $\text{Ba}(\text{OH})_2$.

The student adds 25.0 cm^3 of the $\text{Ba}(\text{OH})_2(\text{aq})$ solution to a conical flask.

The student titrates this solution by adding $0.160 \text{ mol dm}^{-3} \text{ HNO}_3(\text{aq})$ from the burette.

The equation is shown below.



The student repeats the titration until concordant titres are obtained.

The mean titre of $0.160 \text{ mol dm}^{-3} \text{ HNO}_3(\text{aq})$ is 26.75 cm^3 .

- (i) What is meant by concordant titres?

[1]

(ii) Calculate the concentration, in mol dm^{-3} , of the $\text{Ba(OH)}_2(\text{aq})$ solution.

**concentration
of $\text{Ba(OH)}_2(\text{aq})$ = _____ mol dm^{-3} [3]**

(c) A student plans to prepare a solution of $\text{Ba}(\text{OH})_2$ from barium by two different reaction routes.

Outline 2 reaction routes for preparing a solution of $\text{Ba}(\text{OH})_2$ from barium in the laboratory.

Include relevant equations.

[4]

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24 This question is about making ammonia, NH₃.

(a) Ammonia is manufactured by reacting nitrogen with hydrogen:



Standard enthalpy changes of combustion, $\Delta_c H^\ominus$, are given in the table.

Substance	$\Delta_c H^\ominus / \text{kJ mol}^{-1}$
N₂(g)	+180
H₂(g)	−286
NH₃(g)	−293

Calculate the standard enthalpy change of formation, $\Delta_f H^\ominus$, for $\text{NH}_3(\text{g})$.

$\Delta_f H^\ominus$ for $\text{NH}_3(\text{g})$ = _____ kJ mol^{-1} [3]

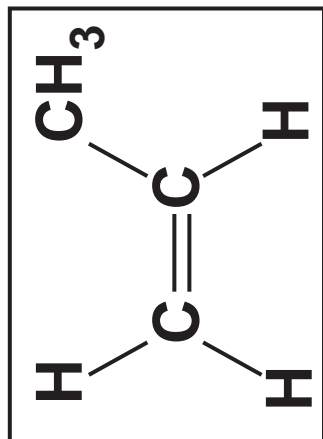
(b) The industrial manufacture of NH_3 from N_2 and H_2 is carried out at an increased temperature and in the presence of a catalyst.

Explain, using Boltzmann distributions, why increasing the temperature and using a catalyst both increase the reaction rate. [5]

25 This question is about unsaturated hydrocarbons.

(a) Two reactions of propene are shown opposite.

In the boxes, show the structures of the organic products of the reactions. [3]



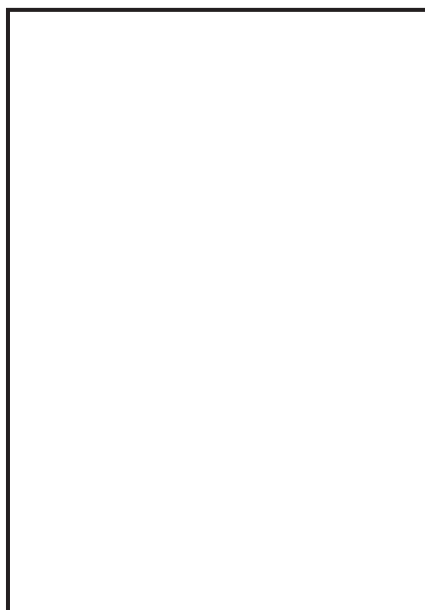
polymerisation



ONE repeat unit



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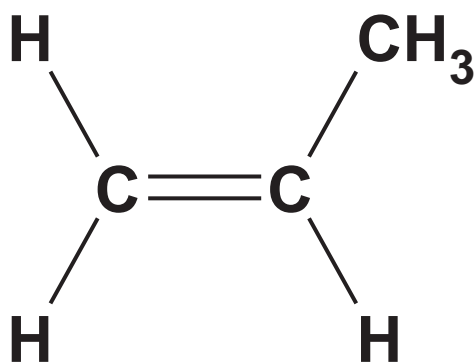


(b) Propene also reacts with bromine.

Outline the mechanism for the reaction of propene with bromine, Br_2 .

The structure of propene has been provided.

Show curly arrows, relevant dipoles and product(s). [4]



(c) The 'alkynes' is a homologous series of hydrocarbons.

The table shows three alkynes.

Alkyne	Structural formula	Molecular formula
ethyne	$\text{HC}\equiv\text{CH}$	C_2H_2
propyne	$\text{CH}_3\text{C}\equiv\text{CH}$	C_3H_4
but-1-yne	$\text{CH}_3\text{CH}_2\text{C}\equiv\text{CH}$	C_4H_6

(i) Explain what is meant by the term: HOMOLOGOUS SERIES.

_____ **[2]**

(ii) Suggest the general formula of the alkynes.

_____ **[1]**

- (iii) Propyne reacts with bromine to form a saturated compound.**

Write an equation for the reaction in the space below, showing the structure of the organic product. [2]

- (iv) But-1-yne is a structural isomer of C_4H_6 .**

Draw the structures of 2 other structural isomers of C_4H_6 in the boxes below.

--	--

[2]

(v) Draw the structure of 2,5-dimethylhept-3-yne. Use the space below. [1]

END OF QUESTION PAPER

ADDITIONAL ANSWER SPACE

If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margin(s).

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[illegible]



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